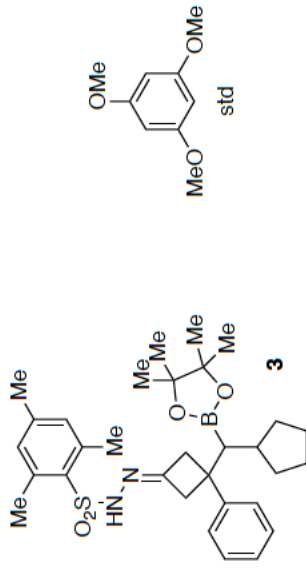


qNMR



$m_{std} = 30.4 \text{ mg}$;
 $MW_{std} = 168.19 \text{ g/mol}$
 $m_3 = 37.3 \text{ mg}$;
 $MW_3 = 550.57 \text{ g/mol}$
molar ratio = $2.24/6 = 0.37$;
 $P_{std} = 99\%$

Thus: $w_3\% = (30.4 \cdot 550.57 \cdot 0.37 \cdot 0.99) / (37.3 \cdot 168.19) \cdot 100$
 $= 97.7\%$

